

CLASS VII – SCIENCE – CHAPTER 04

ACIDS, BASES AND SALTS

Name:

Date:

Q01. Which one is an organic acid?

- (a). Sulphuric acid (b). Hydrochloric acid (c). Nitric acid (d). Lactic acid

Q02. Acid turns blue litmus

- (a). Red (b). Orange (c). Pink (d). green

Q03. Which of the following is a strong acid?

- (a). Citric acid (b). Acetic acid (c). Malic acid (d). Sulphuric acid

Q04. Which gas is released when acid reacts with metals?

- (a). Hydrogen (b). Oxygen (c). Nitrogen (d). Carbon dioxide

Q05. Which of the following is a natural indicator

- (a). Methyl orange (b). Phenolphthalein (c). Turmeric (d). Oxalic acid.

Q06. In case of indigestion, we use

- (a). Antacids (b). Antipyretic (c). Antibiotic (d). Alcohols

Q07. All acids contain

- (a). Oxygen (b). Nitrogen (c). Carbon (d). Hydrogen

Q08. Which of the following is a strong acid?

- (a). Nitric acid (b). Citric acid (c). Tartaric acid (d). Acetic acid

Q09. What is the common name of sodium carbonate?

- (a). Caustic soda (b). Baking soda (c). Phenol (d). Blue vitriol

Q10. Which one of the following metals does not release hydrogen gas with acids?

- (a). Iron (b). Copper (c). Zinc (d). Magnesium

Q11. Sodium chloride turns

- (a). Blue litmus red (b). Red litmus blue
(c). Methyl orange yellow (d). No change in colour

Q12. Reaction between acid and base to form salt is called

- (a). Combination reaction (b). Neutralization reaction
(c). Decomposition reaction (d). Addition reaction

Q01. Match the following

Column A

- (a). Hydrochloric acid
- (b). Carbonic acid
- (c). Sodium hydroxide
- (d). Turmeric
- (e). Blue litmus

Column B

- i. Natural indicator
- ii. Turns red in acid
- iii. Strong acid
- iv. Organic acid
- v. Strong base

Column A

- (a). Turns blue litmus red
- (b). Turns red litmus blue
- (c). Reaction between acid and base
- (d). Gas released when acid react with metal
- (e). Gas turns lime water milky

Column B

- i. Carbon dioxide
- ii. Hydrogen
- iii. Acids
- iv. Bases
- v. Salts

Column A

- (a). Sulphuric acid
- (b). Sodium hydroxide
- (c). Citric acid
- (d). Calcium hydroxide
- (e). Sodium chloride

Column B

- i. Neutral salt
- ii. Weak base
- iii. Strong acid
- iv. Weak acid
- v. Strong base

Q02. Write T for true and F for false statement.

- (a). Sulphuric acid is a strong acid.
- (b). Sodium chloride is a base.
- (c). Litmus is a natural indicator.
- (d). Water soluble bases are called alkalies.
- (e). Bases are non-corrosive in nature

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Q03. Fill in the blanks

- (a). When acid and base react together a new compound called _____ is formed.
- (b). All acids contain _____.
- (c). Metal on heating or burning produce _____ of metals.
- (d). Turmeric is a _____ indicator.
- (e). Minerals acids are obtained from _____.
- (f). The sour things we eat contain _____.
- (g). Ammonium hydroxide is a _____.
- (h). Milk contains _____.
- (i). Litmus is extracted from _____.
- (j). An acid contains more _____ ions.

Q04. What are indicators? Give two examples.

Q05. What is neutralization reaction?

Q06. What are alkalis? Give two examples.

Q07. What is difference between concentrated and dilute acid?

Q08. Why acid solutions are good conductor of electricity?

Q09. Write the properties of bases?

Q10. What is acid rain? How it happens?

Q11. Why curd and sour substance should not be stored in metal containers?

Q12. Differentiate between strong and weak acids with example?

Q13. Classify the following as organic and mineral acids?

Hydrochloric acid, Citric acid, Malic acid, Nitric acid, Oxalic acid, Sulphuric acid.

Q14. Write the colour change when HCl and NaOH is added to

- (a). Blue litmus
- (b). Red litmus
- (c). Methyl orange
- (d). Phenolphthalein